

Mark Scheme (Results)

October 2024

Pearson Edexcel International Advanced Subsidiary Level In Chemistry (WCH13) Paper 01: Practical Skills in Chemistry I

Question Number	Answer		Additional G	uidance	Mark
1(a)(i)	An answer that makes reference to the following points:	Example of completed table:			(4)
	• 6 points correct scores 4		Identity of cation	Colour of flame test	Expert
	5 o points correct scores 1	х	Ba <sup>2+</sup>	(apple) green	Empere
	• 4-5 points correct scores 3	Y	Mg <sup>2+</sup>	no colour	
	• 3 points correct scores 2	z	Ca <sup>2+</sup>	yellow-red / brick red	
			Sr <sup>2+</sup>	(crimson) red	
	• 2 points correct scores 1  TE on correct flame colours for incorrect <b>Group II</b> meta				
		Penalise names and incorrect charges once only (allow TE for subsequent flame colours – max 2 points) Allow Be <sup>2+</sup> for Y – no colour for flame test Allow colourless for no colour Allow orange-red for Ca <sup>2+</sup> (not orange alone) Allow scarlet for Sr <sup>2+</sup>			
			criptors on colours e.g.	pale	
		Ignore nam	es/formulae of salts		
		_	e symbols even if incor	rect im or white for magnesium	
			ard red for calcium	an or miner for magnesium	

Question Number	Answer		Additional Guidance	Mark
<b>1(a)(ii)</b>	A description that makes reference to the following points:			(3)
	• add (dilute) nitric acid / HNO <sub>3</sub> and silver nitrate /AgNO <sub>3</sub> (solution)	(1)	Allow acidified silver nitrate Do not award M1 if extra reagents are added	Expert
	• (chloride will form a) white precipitate	(1)	Accept solid, ppte., ppt etc. Ignore incorrect formulae Do not award chlorine ion Do not award M2 if silver chloride or chloride ions / HCl are added in M1	
	• precipitate will dissolve in <b>dilute</b> (aqueous) ammonia/NH <sub>3</sub> /NH <sub>4</sub> OH solution	(1)	Allow chloride dissolves Allow dilute (aqueous) ammonia hydroxide Ignore concentrated ammonia  Marks are independent but M3 must be after M1	

Question Number	Answer	Additional Guidance	Mark
1(b)(i)	A description that makes reference to the following points:	M2 and M3 can be shown in a labelled diagram, but if marks are awarded then diagram and text must not contradict.	(4) Expert
	• heat the same number of moles of each carbonate (in a test tube over a Bunsen burner)	Allow same mass/amount Allow spirit burner Do not award cotton wool Ignore equal test tubes Do not award same concentration of carbonates	
	• (clamp the test tube so it remains at a) fixed height over the same flame (1	Allow test tube is fixed to stand above a heat source Allow keep heat intensity the same Ignore tripod	
	• connect the test tube to a delivery tube (in)to limewater (1	Ignore gas syringes and collection over water Ignore thermometers	
	• (observe the limewater and) record the time for a precipitate to first appear	Allow cloudiness for precipitate	

<b>Question Number</b>	Answer	Additional Guidance	Mark
1(b)(ii)	An answer that makes reference to the following point:  • a Bunsen flame may not reach the required temperature (to decompose the carbonate)	Allow school equipment will not reach the required temperature (allow examples of equipment) Ignore school laboratory cannot transfer enough thermal	(1) Expert
		energy/heat	

(Total for Question 1 = 12 marks)

/Question Number	Answer		Additional Guidance	Mark
2(a)(i)	• two lines of best fit	(1)	Example of a completed graph  27 26 25 24  Temperature / °C 22 21 20 19 18 17 0 5 10 15 20 25 30 35 40 45 50  Volume of hydrochloric acid / cm³	(3) Expert
			NB lines do not need to be extrapolated beyond the intercept Comment: Allow the right-hand straight line to be a curve but no M1 for a straight line on the left-hand side	
	• reading of volume at intersection	(1)	22.0 (cm <sup>3</sup> ) Accept values between 20.0 and 23.0	
	calculation of maximum temperature change	(1)	(25.2 - 17.5 =) 7.7 Allow values between 7.5 and 7.9	

Question Number	Answer	Additional Guidance	Mark
<b>2(a)(ii)</b>	<ul> <li>moles of hydrochloric acid</li> </ul>		(2)
		Allow answers in the range $0.040 - 0.046$ TE on $2(a)(i)$ Ignore SF	Expert
	• concentration of sodium hydroxide (	0.044 ÷ 0.025 = 1.76 (mol dm <sup>-3</sup> ) Allow answers in the range 1.60 – 1.84 TE from M1 to M2 Ignore SF If given, units must be correct	

Question Number	Answer		Additional Guidance	Mark
<b>2(a)(iii)</b>	<ul> <li>stating or using the equation</li> </ul>	(1)	Example of calculation $\Delta H = (-)mc\Delta T$	(4)
				Expert
	substitution into the equation	(1)	$-(22 + 25) \times 4.18 \times 7.7 = (-)1512.7 \text{ (J)}$ Allow answers in the range $1410 - 1593$ TE on $2(a)(i)$ or answer $-17.5$	
	• calculation of enthalpy per mole	(1)	$1512.7 \div 0.044 = (-)34380 \text{ (J mol}^{-1})$ Allow answers in the range $30652 - 36200$ TE on 2(a)(ii) and M1	
	• conversion to kJ mol <sup>-1</sup>	(1)	24.4 (kl.m.o.1-1)	
	and negative sign	(1)	-34.4 (kJ mol <sup>-1</sup> ) Allow answers in the range -32.7 to -36.2 TE on M2	
			Ignore SF except 1SF TE throughout Correct answer with some working scores 4	

Question Number	Answer		Additional Guidance	Mark
2(b)	An answer that makes reference to two of the following points:			(2)
	• heat loss (to the environment)	(1)	Ignore non-standard conditions	Expert
	• the total volume increases so the temperature change is less even though the same amount of heat is evolved	(1)		
	<ul> <li>the solution is added in (small) portions         OR         the solution is added too slowly</li> </ul>	(1)	Allow incomplete mixing / not stirred	
			Ignore heat capacity/ density are different to water	
			Ignore incomplete reaction, evaporation and impurities	
			Do not award incomplete combustion (lose 1 mark)	

<b>Question Number</b>	Answer	Additional Guidance	Mark
2(c)(i)	An answer that makes reference to the following point:		(1)
	(restart the experiment and) add the estimated volume to the end-point in one portion	Allow add acid in smaller portions (near the endpoint) Allow acid in smaller intervals  Ignore dilute acid	Expert

<b>Question Number</b>	Answer	Additional Guidance	Mark
2(c)(ii)	An answer that makes reference to the following point:		(1)
	add insulation / lid to reduce heat loss     OR     add acid quickly  OR     add acid in larger portions	Ignore draught shield/shielding  Allow increased mixing/swirling Ignore digital thermometer  Ignore increase concentration	Expert

(Total for Question 2 = 13 marks)

Question Number	Answer		Additional Guidance	Mark
3(a)	An answer that makes reference to the following points:		Example of a diagram:	(3)
	An answer that makes reference to the following points:  • condenser (labelled with name) and arrows to indicate the flow of water  • flask (not conical) connected to diagonal condenser with no gaps and unsealed collecting vessel  • thermometer with bulb in correct position, anti-bumping granules and heat source / just arrow	(1) (1) (1)	Example of a diagram:  thermometer  impure product  anti-bumping granules  heat  Thermometer should end within the two red lines Ignore lines across the joints in the apparatus Ignore fractionating columns Entrance and exit of the condenser must be open	(3) Expert
			Please give benefit of doubt on unlabelled anti- bumping granules Allow water in/out labels in place of arrows Allow water in/out labels that contradict direction of arrows Allow attempts at 3D drawing Allow water bath/electric heater Do not award fire for heat	

Question Number	Answer	Additional Guidance	Mark
3(b)(i)	An answer that makes reference to the following point:		(1)
	• there are no C=C bonds (in the product)	Accept the product is saturated / not unsaturated Allow the product is not an alkene Do not award "no double bonds" alone	Graduate

<b>Question Number</b>	Answer	Additional Guidance	Mark
<b>3(b)(ii)</b>	An answer that makes reference to the following point:		(1)
	(Fehling's test shows the final product is) not an aldehyde	Allow does not contain -CHO Allow reflux does not produce aldehydes Ignore speculation about other compounds Ignore comments about oxidation	Graduate

Answer		Additional Guidance	Mark
An answer that makes reference to four of the following points:			(4)
• (original liquid) was not a ketone (as it was oxidised)	(1)	Allow not a ketone as there was a colour change (or converse)	Expert
• (original liquid) has been oxidised	(1)	Allow not tertiary alcohol	
a ketone would (only) be the final product if the purified liquid was a <b>secondary</b> alcohol	(1)	Allow 2° alcohol	
• the original liquid could have been a <b>primary</b> alcohol / aldehyde	(1)		
(the final product could be a) carboxylic acid	(1)	Allow "not enough evidence to be sure of identity of final product" Allow names of specific carboxylic acids Ignore carbonic acid  Ignore all confirmatory tests and colour changes (even if incorrect) Ignore references to tertiary  All marks are independent	
	An answer that makes reference to four of the following points:  • (original liquid) was not a ketone (as it was oxidised)  • (original liquid) has been oxidised  • a ketone would (only) be the final product if the purified liquid was a <b>secondary</b> alcohol  • the original liquid could have been a <b>primary</b> alcohol / aldehyde	An answer that makes reference to four of the following points:  • (original liquid) was not a ketone (as it was oxidised)  • (original liquid) has been oxidised  • (original liquid) has been oxidised  • a ketone would (only) be the final product if the purified liquid was a secondary alcohol  • the original liquid could have been a primary alcohol / aldehyde	An answer that makes reference to four of the following points:  • (original liquid) was not a ketone (as it was oxidised)  • (original liquid) has been oxidised  • (original liquid) has been oxidised  • a ketone would (only) be the final product if the purified liquid was a secondary alcohol  • the original liquid could have been a primary alcohol / aldehyde  • (the final product could be a) carboxylic acid  (1) Allow not tertiary alcohol  (1) Allow 2° alcohol  (1) Allow "not enough evidence to be sure of identity of final product"  Allow not a ketone as there was a colour change (or converse)  (1) Allow not tertiary alcohol  (1) Allow 2° alcohol  (1) In allow "not enough evidence to be sure of identity of final product"  Allow and a ketone as there was a colour change (or converse)

Question Number	Answer		Additional Guidance	Mark
3(c)(ii)	An explanation that makes reference to the following points:			(2)
	• <b>boiling</b> temperature (from the distillation)	(1)	Accept boiling point	Expert
	• can be compared to published values	(1)	Allow data book etc.	

(Total for Question 3 = 11 marks)

Question Number	Answer	Additional Guidance	Mark
4(a)	An answer that makes reference to the following points:	Allow units in brackets throughout	(2) Expert
	• x-axis label: "Time / s" (1)	Allow seconds for s	-
	• y-axis label: "Volume (of H <sub>2</sub> O <sub>2</sub> and foam) / cm <sup>3</sup> " (1)	Allow height /cm	
		Allow 1 mark for both labels in correct locations without units	

<b>Question Number</b>	Answer	Additional Guidance	Mark
4(b)(i)	An answer that makes reference to the following point:		(1)
	• the line (for PbO <sub>2</sub> ) has the steepest gradient	Accept (throughout the experiment) for any given time the volume for PbO <sub>2</sub> is higher Allow steepest slope/line/graph Allow largest volume in the smallest time Allow specific volume reached before the others  Ignore just fastest / fastest rate	Expert

Question Number	Answer		Additional Guidance	Mark
4(b)(ii)	An answer that makes reference to the following points:		Example of a tangent	(3)
	• tangent drawn at steepest point / use of data points from the graph above 36 cm <sup>3</sup>	(1)	45 40 35 30 25 20 0 30 60 90 120 150	Expert
			Do not award M1 if volume used is bigger than 30 unless a tangent is drawn $30 \text{ cm}^3 \div 50 \text{ s}$	
	<ul><li>rate calculated</li><li>correct units</li></ul>	(1)	0.6(0) (Allow 0.52 to 0.68) M2 dependent on M1  cm³ s⁻¹ Accept cm³ / s Allow dm³s⁻¹ with value conversion TE from (a) M3 is independent of M1 and M2	

<b>Question Number</b>	Answer	Additional Guidance	Mark
4(b)(iii)	An answer that makes reference to the following point:		(1)
	• that part of the graph is extrapolated from the earlier points / from 60 s	Allow the height of the foam is difficult to measure (accurately) Allow only two readings were taken (to find the shape of the graph)  Allow more readings need to be taken after 60 s Allow extrapolation is difficult Allow a description of extrapolation	Expert

Question Number	Answer	Additional Guidance	Mark
4(c)(i)	An answer that makes reference to the following point:		(1)
	• the catalyst may not be fully mixed with the (H <sub>2</sub> O <sub>2</sub> ) solution at first	Allow because there is an induction period / the catalyst doesn't work effectively straight away Allow the contact area for the reactant and catalyst increase during the reaction Allow the temperature increases, increasing the rate of reaction Allow the reaction is exothermic  Ignore parallax errors	Expert

Question Number	Answer		Additional Guidance	Mark
4(c)(ii)	An answer that makes reference to the following points:		Example of a curve:	(2)
	initial similar shape to previous graph	(1)	i.e. slow initial rate that increases Do not award if line starts above the origin Allow multiple lines of the correct shape	Expert
	• convex curve to plateau at maximum height	(1)	Allow a dip after the plateau has been reached Do not award M2 for a line that returns to the axis	

Question Number	Answer	Additional Guidance	Mark
4(d)(i)	An answer that makes reference to the following point:  • corrosive symbol drawn	Either one is acceptable for the mark	(1) Expert
		(container, liquid dripping and a hand/block underneath showing some damage / zig zags)	

Question Number	Answer	Additional Guidance	Mark
4(d)(ii)	An answer that makes reference to the following points:		(1)
	• gloves	Ignore use a fume cupboard Ignore wear a mask Ignore tongs Ignore modifiers e.g. insulating gloves Ignore comments about volumes and other equipment  Do not award "not touching the liquid"	Graduate

Question Number	Answer	Additional Guidance	Mark
4(d)(iii)	An answer that makes reference to the following point:		(1)
	• dilute the peroxide/H <sub>2</sub> O <sub>2</sub>	Allow lower/smaller concentration Ignore decrease amount/volume/mass	Graduate

Question Number	Answer	Additional Guidance	Mark
4(d)(iv)	An answer that makes reference to the following point:		(1)
	• a suitable reason	E.g. To stop the peroxide damaging the table / surface So that spills can be easily cleared up In case the foam spills over the top  Ignore corrosive Ignore stop contact with skin Ignore references to stability / falling over Do not award references to acid	Expert

(Total for Question 4 = 14 marks)

**TOTAL FOR PAPER = 50 MARKS**